

Ignition delay time model based on a deep neural network

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1 Introduction

The ignition delay time (IDT) is one of basic properties describing any flammable mixture and it is very important in process safety management. IDT is usually modeled with detailed reaction mechanisms (DRMs), however those calculations might be very time consuming and DRMs are still being developed and refined. The authors of the recent paper accumulated around 1800 IDTs from shock tube experiments for various C1–C7 hydrocarbon–O₂–Ar mixtures. This amount of points becomes sufficient for deployment of one of Machine Learning algorithms - a deep neural network (DNN). Machine Learning is widely used in numerous aspects of life, for instance: self-driving cars, handwriting recognition, anti-spam filtering, web search and rating systems. Now it is becoming more popular in science as well. A DNN is based on an artificial neural network which is inspired by the biological neural networks. The biggest advantage of DNNs is high predictive power and flexibility of application. According to authors' best knowledge IDTs have not been yet modeled with a DNN. Hence, the goal of the recent paper is to introduce a new ignition delay time model based on a DNN technique.

2 Model

A large data set of ignition delay times is collected for C1–C7 hydrocarbon-oxygen-argon mixtures (1789 points in total) from shock tube experiments [1–28]. The data set can be briefly summarized as follows:

- the collected IDT range is from 2.6 μ s up to 2.11 ms;
- C1–C7 hydrocarbons are methane, ethane, ethene, acetylene, propane, propene, *iso*- and *n*-butane, *I*- and *iso*-butene, *n*-, *neo*- and *iso*-pentane, *I*-pentene, *n*-hexane, *I*-hexene and *n*-heptane;
- temperature range is 1022–2596 K, pressure 0.3–267 bar, EQR 0.06–4.0, argon molar fraction 54.5%–99.93%.

The DNN model is developed based on following assumptions:

- the whole dataset of 1789 examples is divided into train and test sets in a proportion of 80/20;
- the number of inputs is equal to 7 (fuel type, fuel molar fraction, oxygen molar fraction, argon molar fraction, initial pressure (bar), initial temperature (K), inverse of initial temperature (1/K), one can note that higher number of inputs is desirable from feature engineering point of view);
- the output is ignition delay time (μs);
- activation function for all hidden layers is hyperbolic tangent;
- identity function is used for the output layer.

The deployment of the model structure presented in Figure 1 (the number of hidden layers, the number of nodes in each layer and the L2 regularization parameter [29]) is described in detail in the paper of Malik et al. [30].

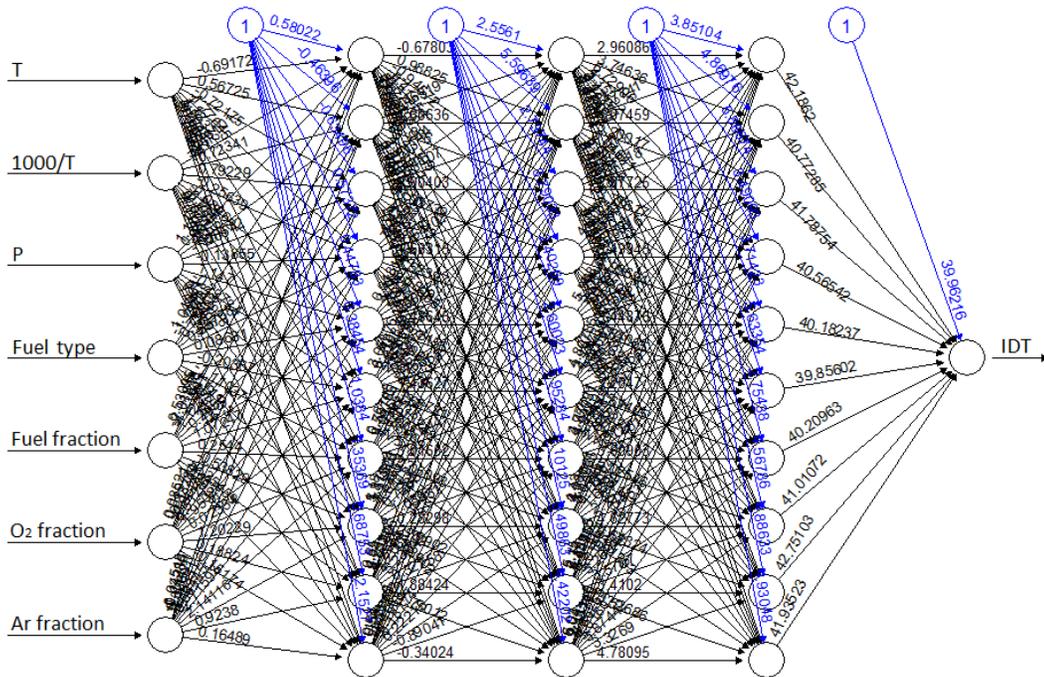


Figure 1: Deep neural network structure used in the present study

3 Results

The performance of the DNN model is measured by Mean Absolute Error (MAE) and the Pearson coefficient (CORR), which measures linear correlation between two sets of values. CORR is defined as follows:

$$CORR(x, y) = \frac{\sum_{i=1}^n (x_i - \bar{x})(y_i - \bar{y})}{\sqrt{\sum_{i=1}^n (x_i - \bar{x})^2} \sqrt{\sum_{i=1}^n (y_i - \bar{y})^2}} \quad (1)$$

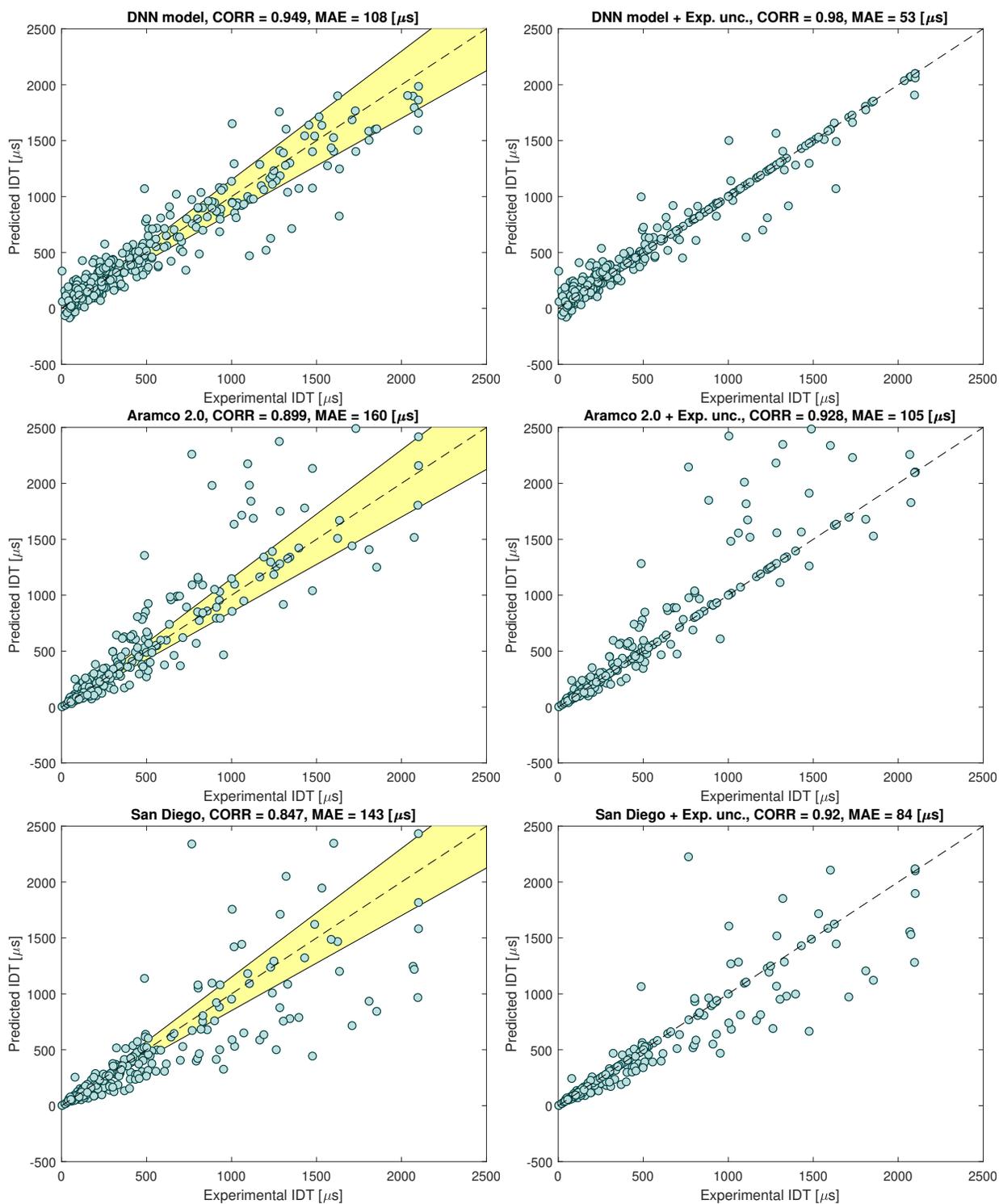


Figure 2: Predicted IDT vs. experimental IDT. Comparison of DNN model to Aramco 2.0 and San Diego 2014. Left: Experimental uncertainty is marked by the yellow area. Right: Experimental uncertainty is taken into account in CORR and MAE calculations.

where x and y are vectors of experimental and predicted ignition delay times, n is the sample size (size of both samples is equal).

In Figure 2 the DNN performance on a test set is presented and compared to two well-known DRMs: Aramco 2.0 [5,28,31–35] and San Diego 2014 [36]. In the left column of Figure 2 experimental uncertainty range (assumed as $IDT \pm 15\%$) is marked by yellow area. Predictions belonging to the *area* are considered as perfect match to experiment. In the right part of Figure 2 experimental uncertainty is taken into account while calculating CORR and MAE. Additionally, in Table 1 the DNN model is compared with NUIG n-Heptane [37] and GRI-mech 3.0 [38]. The DNN model reaches higher CORR and lower MAE than the *best* DRMs' (NUIG n-Heptane: MAE = 85 μ s, CORR = 0.933, the DNN: MAE = 53 μ s, CORR = 0.98, Table 1).

Table 1: Comparison of DNN model to DRMs.

Metric	DNN	Aramco 2.0	NUIG n-Heptane	GRI-mech 3.0	San Diego 2014
<i>without experimental uncertainties</i>					
CORR	0.949	0.899	0.902	0.645	0.847
MAE	108	160	136	599	143
<i>with experimental uncertainties</i>					
CORR	0.980	0.928	0.933	0.648	0.920
MAE	53	105	85	547	84

4 Conclusion

The DNN model for IDT of C1–C7 hydrocarbon-oxygen-argon mixtures is developed. It results in lower error than the *best* DRM (NUIG n-Heptane) and much shorter computational time (DNN: less than 1 ms, NUIG n-Heptane: 2 min). The model can be easily extended to nitrogen diluted mixtures and to new experimental data as published. In the future work the model for ignition delay times from rapid compression machines and for mixtures diluted with nitrogen will be developed.

References

- [1] C. J. Aul, W. K. Metcalfe, S. M. Burke, H. J. Curran, and E. L. Petersen, "Ignition and kinetic modeling of methane and ethane fuel blends with oxygen: A design of experiments approach," *Combustion and Flame*, vol. 160, no. 7, pp. 1153–1167, 2013. [Online]. Available: <http://dx.doi.org/10.1016/j.combustflame.2013.01.019>
- [2] J. A. Baker and G. B. Skinner, "Shock-tube studies on the ignition of ethylene-oxygen-argon mixtures," *Combustion and Flame*, vol. 19, no. 3, pp. 347–350, 1972.
- [3] C. J. Brown and G. O. Thomas, "Experimental studies of shock-induced ignition and transition to detonation in ethylene and propane mixtures," *Combustion and Flame*, vol. 117, no. 4, pp. 861–870, 1999.

- [4] J. Bugler, K. P. Somers, E. J. Silke, and H. J. Curran, "Revisiting the kinetics and thermodynamics of the low-temperature oxidation pathways of alkanes: a case study of the three pentane isomers," *The Journal of Physical Chemistry A*, vol. 119, no. 28, pp. 7510–7527, 2015.
- [5] S. M. Burke, U. Burke, R. Mc Donagh, O. Mathieu, I. Osorio, C. Keesee, A. Morones, E. L. Petersen, W. Wang, T. A. DeVerter, M. A. Oehlschlaeger, B. Rhodes, R. K. Hanson, D. F. Davidson, B. W. Weber, C. J. Sung, J. Santner, Y. Ju, F. M. Haas, F. L. Dryer, E. N. Volkov, E. J. K. Nilsson, A. A. Konnov, M. Alrefae, F. Khaled, A. Farooq, P. Dirrenberger, P. A. Glaude, F. Battin-Leclerc, and H. J. Curran, "An experimental and modeling study of propene oxidation. Part 2: Ignition delay time and flame speed measurements," *Combustion and Flame*, vol. 162, no. 2, pp. 296–314, 2015.
- [6] Y. Cheng, E. Hu, F. Deng, F. Yang, Y. Zhang, C. Tang, and Z. Huang, "Experimental and kinetic comparative study on ignition characteristics of 1-pentene and n-pentane," *Fuel*, vol. 172, pp. 263–272, 2016. [Online]. Available: <http://dx.doi.org/10.1016/j.fuel.2016.01.008>
- [7] D. F. Davidson and R. K. Hanson, "Interpreting shock tube ignition data," *International Journal of Chemical Kinetics*, vol. 36, no. 9, pp. 510–523, 2004.
- [8] —, "Fundamental kinetics database utilizing shock tube measurements," ... *Engineering Department, Stanford* ... , vol. 1, no. January, pp. 1–75, 2005. [Online]. Available: <http://hanson.stanford.edu/researchReports/kinetics/Fundamental{ }Kinetics{ }Database{ }Volume{ }1.pdf>
- [9] —, "Fundamental kinetics database utilizing shock tube measurements," ... *Engineering Department, Stanford* ... , vol. 4, no. January, pp. 1–144, 2014. [Online]. Available: <http://hanson.stanford.edu/researchReports/kinetics/Fundamental{ }Kinetics{ }Database{ }Volume{ }1.pdf>
- [10] B. Eiteneer and M. Frenklach, "Experimental and modeling study of shock-tube oxidation of acetylene," *International Journal of Chemical Kinetics*, vol. 35, no. 9, pp. 391–414, 2003.
- [11] P. Frank and M. Braun-Unkhoff, "A shock tube study on the reaction: $\text{C}_3\text{H}_3 + \text{C}_2\text{H}_5 \rightarrow \text{C}_2\text{H}_4 + \text{H}$," *Shock tubes and waves*, pp. 379–385, 1988.
- [12] Y. Hidaka, "Shock tube and modeling study of the ignition chemistry of small hydrocarbons," *J. Mol. Sci.*, vol. 23, pp. 141–153, 1982.
- [13] Y. Hidaka, K. Hattori, T. Okuno, K. Inami, T. Abe, and T. Koike, "Shock tube and modeling study of acetylene pyrolysis and oxidation," *Combustion and Flame*, vol. 107, pp. 401–417, 1996.
- [14] E. Hu, Y. Chen, Z. Zhang, X. Li, Y. Cheng, and Z. Huang, "Experimental Study on Ethane Ignition Delay Times and Evaluation of Chemical Kinetic Models," *Energy & Fuels*, vol. 29, no. 7, pp. 4557–4566, 2015. [Online]. Available: <http://pubs.acs.org/doi/abs/10.1021/acs.energyfuels.5b00462>
- [15] E. Hu, Z. Gao, Y. Liu, G. Yin, and Z. Huang, "Experimental and modeling study on ignition delay times of dimethoxy methane/n-heptane blends," *Fuel*, vol. 189, pp. 350–357, feb 2017. [Online]. Available: <http://linkinghub.elsevier.com/retrieve/pii/S0016236116310699>
- [16] N. Lamoureux, C. E. Paillard, and V. Vaslier, "Low hydrocarbon mixtures ignition delay times investigation behind reflected shock waves," *Shock Waves*, vol. 11, no. 4, pp. 309–322, 2002.

- [17] A. Lifshitz, K. Scheller, A. Burcat, and G. B. Skinner, "Shock-tube investigation of ignition in methane-oxygen-argon mixtures.pdf," *Combustion and Flame*, vol. 321, no. 3, pp. 311–321, 1971.
- [18] X. Man, C. Tang, L. Wei, L. Pan, and Z. Huang, "Measurements and kinetic study on ignition delay times of propane/hydrogen in argon diluted oxygen," *International Journal of Hydrogen Energy*, vol. 38, no. 5, pp. 2523–2530, 2013. [Online]. Available: <http://dx.doi.org/10.1016/j.ijhydene.2012.12.020>
- [19] L. Pan, Y. J. Zhang, J. X. Zhang, Z. M. Tian, and Z. H. Huang, "Shock tube and kinetic study of C₂H₆/H₂/O₂/Ar mixtures at elevated pressures," *International Journal of Hydrogen Energy*, vol. 39, no. 11, pp. 6024–6033, 2014.
- [20] L. Pan, E. Hu, J. Zhang, Z. Tian, X. Li, and Z. Huang, "A high pressure shock tube study of 1-butene oxidation and its comparison with n-butane and alkenes," *Fuel*, vol. 157, pp. 21–27, 2015. [Online]. Available: <http://dx.doi.org/10.1016/j.fuel.2015.04.062>
- [21] D. J. Seery and C. T. Bowman, "An experimental and analytical study of methane oxidation behind shock waves," *Combustion and Flame*, vol. 14, no. 1, pp. 34–47, 1970.
- [22] L. J. Spadaccini and M. B. Colket, "Ignition delay characteristics of methane fuels," *Progress in Energy and Combustion Science*, vol. 20, no. 5, pp. 431–460, 1994.
- [23] C. Tang, X. Man, L. Wei, L. Pan, and Z. Huang, "Further study on the ignition delay times of propane-hydrogen-oxygen-argon mixtures: Effect of equivalence ratio," *Combustion and Flame*, vol. 160, no. 11, pp. 2283–2290, 2013.
- [24] T. Tsuboi and H. G. Wagner, "Homogeneous thermal oxidation of methane in reflected shock waves," *Symposium (International) on Combustion*, vol. 15, no. 1, pp. 883–890, 1975.
- [25] J. de Vries, J. M. Hall, S. L. Simmons, M. J. A. Rickard, D. M. Kalitan, and E. L. Petersen, "Ethane ignition and oxidation behind reflected shock waves," *Combustion and Flame*, vol. 150, no. 1-2, pp. 137–150, 2007.
- [26] M. Yahyaoui, N. Djebaili-Chaumeix, P. Dagaut, C. E. Paillard, and S. Gail, "Kinetics of 1-hexene oxidation in a JSR and a shock tube: Experimental and modeling study," *Combustion and Flame*, vol. 147, no. 1-2, pp. 67–78, 2006.
- [27] Y. Zhang, Z. Huang, L. Wei, J. Zhang, and C. K. Law, "Experimental and modeling study on ignition delays of lean mixtures of methane, hydrogen, oxygen, and argon at elevated pressures," *Combustion and Flame*, vol. 159, no. 3, pp. 918–931, 2012. [Online]. Available: <http://dx.doi.org/10.1016/j.combustflame.2011.09.010>
- [28] C. W. Zhou, Y. Li, E. O'Connor, K. P. Somers, S. Thion, C. Keesee, O. Mathieu, E. L. Petersen, T. A. DeVerter, M. A. Oehlschlaeger, G. Kukkadapu, C. J. Sung, M. Alrefae, F. Khaled, A. Farooq, P. Dirrenberger, P. A. Glaude, F. Battin-Leclerc, J. Santner, Y. Ju, T. Held, F. M. Haas, F. L. Dryer, and H. J. Curran, "A comprehensive experimental and modeling study of isobutene oxidation," *Combustion and Flame*, vol. 167, pp. 353–379, 2016.
- [29] A. E. Hoerl and R. W. Kennard, "Ridge regression: Biased estimation for nonorthogonal problems," *Technometrics*, vol. 12, no. 1, pp. 55–67, feb 1970.

- [30] K. Malik, M. Żbikowski, and A. Teodorczyk, “Detonation cell size model based on deep neural network for hydrogen, methane and propane mixtures with air and oxygen,” *Nuclear Engineering and Technology*, nov 2018.
- [31] U. Burke, W. K. Metcalfe, S. M. Burke, K. A. Heufer, P. Dagaut, and H. J. Curran, “A detailed chemical kinetic modeling, ignition delay time and jet-stirred reactor study of methanol oxidation,” *Combustion and Flame*, vol. 165, pp. 125–136, mar 2016. [Online]. Available: <http://linkinghub.elsevier.com/retrieve/pii/S0010218015003958>
- [32] S. M. Burke, W. Metcalfe, O. Herbinet, F. Battin-Leclerc, F. M. Haas, J. Santner, F. L. Dryer, and H. J. Curran, “An experimental and modeling study of propene oxidation. Part 1: Speciation measurements in jet-stirred and flow reactors,” *Combustion and Flame*, vol. 161, no. 11, pp. 2765–2784, nov 2014. [Online]. Available: <http://linkinghub.elsevier.com/retrieve/pii/S0010218014001400>
- [33] W. K. Metcalfe, S. M. Burke, S. S. Ahmed, and H. J. Curran, “A Hierarchical and Comparative Kinetic Modeling Study of C1-C2 Hydrocarbon and Oxygenated Fuels,” *International Journal of Chemical Kinetics*, vol. 45, no. 10, pp. 638–675, oct 2013. [Online]. Available: <http://doi.wiley.com/10.1002/kin.20802>
- [34] A. Kéromnès, W. K. Metcalfe, K. A. Heufer, N. Donohoe, A. K. Das, C.-J. Sung, J. Herzler, C. Naumann, P. Griebel, O. Mathieu, M. C. Krejci, E. L. Petersen, W. J. Pitz, and H. J. Curran, “An experimental and detailed chemical kinetic modeling study of hydrogen and syngas mixture oxidation at elevated pressures,” *Combustion and Flame*, vol. 160, no. 6, pp. 995–1011, jun 2013. [Online]. Available: <http://linkinghub.elsevier.com/retrieve/pii/S0010218013000023>
- [35] Y. Li, C.-W. Zhou, K. P. Somers, K. Zhang, and H. J. Curran, “The oxidation of 2-butene: A high pressure ignition delay, kinetic modeling study and reactivity comparison with isobutene and 1-butene,” *Proceedings of the Combustion Institute*, vol. 000, pp. 1–9, 2016. [Online]. Available: <http://linkinghub.elsevier.com/retrieve/pii/S1540748916300529>
- [36] Mechanical and Aerospace Engineering (Combustion Research) - University of California at San Diego, “Chemical-Kinetic Mechanisms for Combustion Applications, San Diego Mechanism web page,” 2018. [Online]. Available: <http://web.eng.ucsd.edu/mae/groups/combustion/mechanism.html>
- [37] K. Zhang, C. Banyon, J. Bugler, H. J. Curran, A. Rodriguez, O. Herbinet, F. Battin-Leclerc, C. B’Chir, and K. A. Heufer, “An updated experimental and kinetic modeling study of n-heptane oxidation,” *Combustion and Flame*, vol. 172, pp. 116–135, 2016. [Online]. Available: <http://dx.doi.org/10.1016/j.combustflame.2016.06.028>
- [38] G. P. Smith, D. M. Golden, M. Frenklach, N. W. Moriarty, B. Eiteneer, C. T. Bowman, R. K. Hanson, S. Song, W. C. Gardiner, V. V. Lissianski, and Z. Qin, “GRI-Mech 3.0,” 2018. [Online]. Available: http://www.me.berkeley.edu/gri_{-}mech/{%}0A